

## Unit 10: Heat of Reactions Practice

Name: \_\_\_\_\_

### Checklist:

- ✓ Is the equation given?
- ✓ Is it balanced?
- ✓ Which element/compound does the problem give a mass for?
- ✓ How many moles of that substance are in the balanced equation? How many grams is that?
- ✓ Set up a proportion or t-chart. What is given? What should you have in a perfect world? We have perfect heat, find the actual heat.

How much heat will be released when 6.44g of sulfur reacts with excess O<sub>2</sub> according to the following equation?



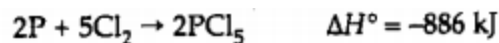
How much heat will be released when 4.72g of carbon reacts with excess O<sub>2</sub> according to the following equation?



How much heat will be absorbed when 38.2g of bromine reacts with excess H<sub>2</sub> according to the following equation?



How much heat will be released when 1.48g of chlorine reacts with excess phosphorus according to the following equation?



How much heat will be released when 4.77g of ethanol (C<sub>2</sub>H<sub>5</sub>OH) reacts with excess O<sub>2</sub> according to the following equation?

